Peer Instruction workshop

He : EOTON :: Gang



5/E Conference on Advancing Teaching and Learning 410 6IE Texas Tech University ERIC MAZUR Lubbock, TX, 26 February 2010

Ready-to-Use

Resources

1 Tichnole

310

814

GIF

2B

716

0

class-

C.C.

Get your clickers ready!



- no ON/OFF button
- only last "click" counts
- display shows recorded answer

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Consider this

Professors A and B teach the same mechanics class at the same college during different semesters. Professor A uses the traditional approach to teaching and lectures. Professor B uses Peer Instruction and students respond to the questions in using clickers. Each class is evaluated using the traditional end-ofsemester questionnaire and using the FCI to measure students' comprehension of mechanics. Both professors are middle-aged and male. The results are as follows.

A: student evaluation: 1.5/5.0; <g> = 0.42 B: student evaluation: 3.7/5.0; <g> = 0.57

Consider this

Professor	Α	В
pedagogy	traditional	PI with clickers
student evaluation	1.5/5.0	3.7/5.0
FCI <g></g>	0.42	0.57

What might account for the large difference in evaluation?

I. professor personality II. technology III. pedagogy

1. I only 2. II only 3. III only 4. II and III5. I, II, and III6. other combination





1. you got engaged



- 1. you got engaged
- 2. no "correct" answer



- 1. you got engaged
- 2. no "correct" answer
- 3. you got engaged



- 1. you got engaged
- 2. no "correct" answer
- 3. you got engaged
- 4. you don't need a correct answer!

Outline

20



Outline

Setting the stage

• Making it happen

• Overcoming barriers

To set stage for successful implementation, I need to...

- convince yourself (and your colleagues)
- set learning goals
- select approaches
- identify resources





Evaluate assessment by comparing

student performance on various kinds of problems

What constitutes a good problem?

On a Saturday afternoon, you pull into a parking lot with unmetered spaces near a shopping area. You circle around, but there are no empty spots. You decide to wait at one end of the lot, where you can see (and command) about 20 spaces.

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Requires:

Assumptions Developing a model Applying that model

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Requires:

Applying a (new) model

On a Saturday afternoon, you pull into a parking lot with unmetered spaces near a shopping area, where people are known to shop, on average, for 2 hours. You circle around, but there are no empty spots. You decide to wait at one end of the lot, where you can see (and command) about 20 spaces.

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Requires:

Using a calculator

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Need to test meaningful skills!

Setting learning goals



Setting learning goals



approach, not content

focus on understanding

Traditional approach to course planning



Traditional approach to course planning



Traditional approach to course planning














How to move information transfer out of classroom?

Approaches for reading:

- Reading quizzes
- Summaries
- Just-in-time-Teaching (www.jitt.org)



Are clickers a required resource?

Flashcards: simple and effective



Flashcards: simple and effective



Meltzer and Mannivanan, South Eastern Louisiana University

Imagine a rope that fits snugly along the equator.



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Suppose the rope is cut and 1 m of rope is inserted between the cut ends. If the rope were to maintain a circular shape, how far off the surface of the Earth would it float?

- 1. the width of a few atoms
- 2. the width of a few hairs
- 3. the height of a curb
- 4. exactly 1 m
- 5. more than 1 m



circumference at equator:

 $2\pi R_{\rm E}$

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new circumference:

 $2\pi R_{\rm E} + 1 \,{\rm m}$

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radius of circle with new circumference:

 $2\pi R = 2\pi R_{\rm E} + 1 \,{\rm m}, \text{ and so } R = R_{\rm E} + \frac{1 \,{\rm m}}{2\pi}.$

You all got fired up!

You all got fired up!

(WITHOUT CLICKERS!)



It's not the technology, but the pedagogy!



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(but clickers do offer advantages)

Outline

Setting the stage

• Making it happen

Overcoming barriers

To make it happen, I need to...

find or develop good questions

know how to manage time









Videos:

- Interactive Teaching DVD
- From questions to concepts



Types of questions

- survey
- discussion
- model testing
- select from list

Which of the following airlines tries to save fuel by suggesting that its passengers use the bathroom before boarding?

- **1. Delta Airlines**
- 2. Lufthansa
- 3. All Nippon Airways
- 4. British Midland Airways
- 5. Air France
- 6. JAL
- 7. Aboriginal Air Services
- 8. Aeroflot
- 9. Are you kidding me? None of the above.

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hole in plate/circumference model

Professor A/B

discussion

airline

fact

fact-recall not engaging

Good conceptual questions (ConcepTests):

- are based on student difficulties
- focus on single concept
- cannot be solved by "plug and chug"
- are clear and concise
- are of manageable difficulty

must adjust level to audience














ConcepTest data



brief presentation























Outline

Setting the stage

Making it happen

What are some potential barriers?

- skepticism
- growing pains
- negative feedback
- limited circle of influence

After changing, things might get *worse* before they get better!





Better understanding leads to *more* — not fewer — questions!

(must recognize confusion as step towards understanding)

Things to do:

- take data
- motivate students
- be prepared for initial adjustments

Why is change so hard?

	"lectures"	PI
coverage	complete	partial
preclass reading	none	cover everything
confusion	little none	substantial
evaluations	known	unknown

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(neither of which is traditionally measured)





Conclusion

active engagement greatly improves learning gains

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technology facilitates active engagement

Conclusion

not just a polling tool, but an engagement tool!

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Let's try it!

The specific heat at constant volume for a monatomic crystal approaches zero at low temperature even though the specific heat for a monatomic gas remains $\frac{3}{2}k$ per atom. Why is this so?

Let's try it!

The specific heat at constant volume for a monatomic crystal approaches zero at low temperature even though the specific heat for a monatomic gas remains $\frac{3}{2}k$ per atom. Why is this so?

- 1. Potential energy doesn't play a role for the monatomic gas, but it does for the crystal.
- 2. The particles are indistinguishable in the gas, but not in the crystal.
- 3. The energy difference between allowed states for the crystal is much larger than it is for the atoms.